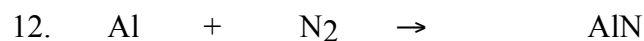
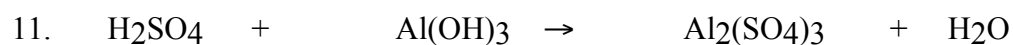
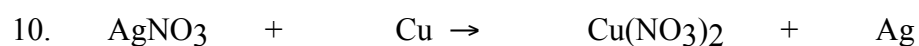
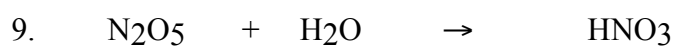
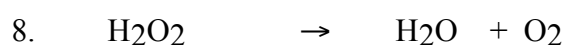
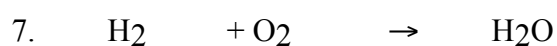
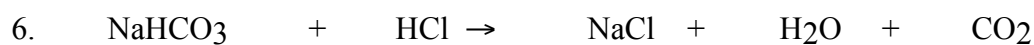
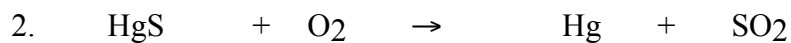
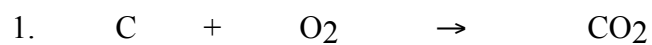
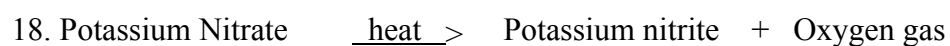
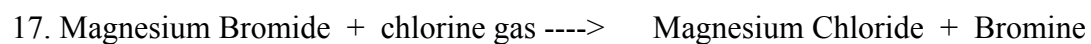
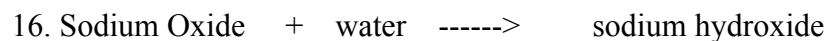
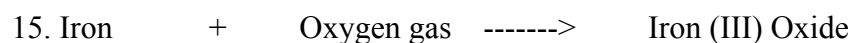


Practice: Balance the following equations



Practice: Write balanced chemical equations for the following



19. zinc + hydrochloric Acid ----> Zinc Chloride + Hydrogen gas

20. calcium oxide + hydrochloric Acid ----> calcium chloride + water

21. Chlorine gas + sodium bromide --> sodium chloride + bromine

22. zinc + Sulfuric acid --> zinc sulfate + hydrogen gas

23. Phosphorus + oxygen gas --> diphosphorus pentoxide

24. tin + oxygen gas --> tin(IV) oxide

25. hydrogen gas + iodine --> hydrogen iodide

26. Solid DiIodine pentoxide and carbon monoxide react to produce elemental Iodine and carbon dioxide.

27. In the presence of MnO₂ (a catalyst), Hydrogen peroxide will decompose into water and oxygen gas

28. Elemental copper & sulfuric acid react to produce aqueous copper (II) sulfate, gaseous sulfur dioxide, water

29. Solid Iron II sulfide reacts with pure oxygen to produce Iron III oxide and sulfur dioxide